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With Answer Key **Worksheet On Balancing Chemical Equations**

When you find difficulty in balancing the equation in the balancing chemical equations worksheet, you can miss it with a fraction of $\frac{1}{2}$ and that will easily balance the equation. But the problem is that you cannot have a fraction for the co-efficient, this is why doubling all

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coefficients will help you balance the equation.

49 Balancing Chemical Equations Worksheets [with Answers]

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your practice. Parts of a Balanced Chemical Equation. Before you start balancing chemical equations, it is important that you become firmly acquainted with the various part of one.

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REVIEW OF WRITING CHEMICAL

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EQUATIONS For each of the following equations write a formula equation. YOU DO NOT HAVE TO BALANCE IT. When applicable, write a net ionic equation.

1. Solutions of sodium fluoride and dilute hydrochloric acid are mixed.
2. A saturated solution of barium hydroxide is mixed with a solution of iron(III) sulfate.
- 3.

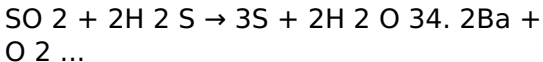
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BALANCING CHEMICAL EQUATIONS - PRACTICE SHEET # 2

Balancing Chemical Equations

Worksheet 2 - ANSWERS
26. $2\text{Mg} + \text{Cl}_2 \rightarrow 2\text{MgCl}_2$
27. $2\text{Ag}_2\text{O} \rightarrow 4\text{Ag} + \text{O}_2$
28. $4\text{K} + 2\text{O}_2 \rightarrow 2\text{K}_2\text{O}$
29. $\text{Cl}_2 + 3\text{F}_2 \rightarrow 2\text{ClF}_3$
30. $\text{SiO}_2 + 2\text{C} \rightarrow \text{Si} + 2\text{CO}$
31. $2\text{NaHCO}_3 \rightarrow \text{Na}_2\text{CO}_3 + \text{CO}_2 + \text{H}_2\text{O}$

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Balancing Chemical Equations Worksheet 1

Balancing Chemical Equations

Worksheet Science Spot Answers.

Answer Key Writing And Balancing

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Balance the following chemical equations.

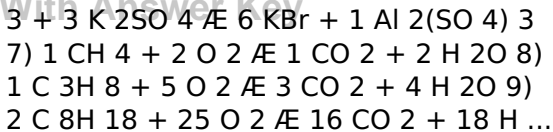
- $\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + \text{H}_2$
- $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{CO}_2$
- $\text{KOH} + \text{H}_3\text{PO}_4 \rightarrow \text{K}_3\text{PO}_4 + \text{H}_2\text{O}$
- $\text{SnO}_2 + \text{H}_2 \rightarrow \text{Sn} + \text{H}_2\text{O}$
- $\text{NH}_3 + \text{O}_2 \rightarrow \text{NO} + \text{H}_2\text{O}$
- $\text{KNO}_3 + \text{H}_2\text{CO}_3 \rightarrow \text{K}_2\text{CO}_3 + \text{HNO}_3$
- $\text{B}_2\text{Br}_6 + \text{HNO}_3 \rightarrow \text{B}(\text{NO}_3)_3 + \text{HBr}$
- $3\text{BF}_3 + \text{Li}_2\text{SO}_3 \rightarrow \text{B}_2(\text{SO}_3)_3 + \text{LiF}$
- $(\text{NH}_4)_3\text{PO}_4 + \text{Pb}(\text{NO}_3)_4 \rightarrow \text{Pb} \dots$

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Balancing Equations: Practice Problems

Balancing Chemical Equations Answer Key Balance the equations below: 1) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ 2) $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$ 3) $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$ 4) $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$ 5) $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$ 6) 2 AlBr_3

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Chapter 7 Worksheet #1 Balancing Chemical Equations

Balancing Chemical Equations - Answer
Key Balance the equations below: 1) 1

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$N_2 + 3 H_2 \rightarrow 2 NH_3$ 2) $2 KClO_3 \rightarrow 2 KCl + 3 O_2$ 3) $2 NaCl + 1 F_2 \rightarrow 2 NaF + 1 Cl_2$
4) $2 H_2 + 1 O_2 \rightarrow 2 H_2O$ 5) $1 Pb(OH)_2 + 2 HCl \rightarrow 2 H_2O + 1 PbCl_2$ 6) $2 AlBr_3 + 3 K_2SO_4 \rightarrow 6 KBr + 1 Al_2(SO_4)_3$ 7) $1 CH_4 + 2 O_2 \rightarrow 1 CO_2 + 2 H_2O$ 8) $1 C_3H_8 + 5 O_2 \rightarrow 3 CO_2 + 4 H_2O$ 9) $2 C_8H_{18} \dots$

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What is the function of the parentheses in chemical equations. Take our photosynthesis example earlier. The

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product $C_6H_{12}O_6$ is a sugar called glucose. When the chemical equation was balanced, this compound appeared as $6(C_6H_{12}O_6)$. This means that there are six of the glucose compound.

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